

CHOOSE THE BEST ANSWER

1. Which one is the molecules of the Iron.  
(a)  $\text{Cl}^-$  (b)  $\text{Na}^+$  (c)  $\text{NH}_4^+$  (d)  $\text{S}^{2-}$
2. What is the atomic number of Ne  
(a) 8 (b) 9 (c) 10 (d) 11
3. How many electrons present in Hydrogen atom  
(a) 1 (b) 2 (c) 3 (d) 4
4. Which one is the lightest element  
(a) He (b) Li (c) Be (d) B
5. What is the meaning of "Halogens"  
(a) Salt (b) Acid (c) Base (d)  $\text{H}_2\text{O}$
6. What is the meaning of "Bivalency"  
(a) 3 (b) 2 (c) 1 (d) 0
7. Poly atomic molecule is  
(a)  $\text{CH}_4$  (b)  $\text{Cl}_2$  (c)  $\text{F}_2$  (d)  $\text{I}_2$
8. The concentration of the solution is directly proportional to  
(a) T (b) P (c) V (d) None of these.  
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9. Ideal gas equation is  
(a)  $PV/nRT$  (b)  $PV=nRT$  (c)  $PV/T$   
(d)  $PV/RT$
10. Which one is non-metal from the following.  
(a) K (b) Na (c) C (d) Li
11. Which one is "metalloid" element  
(a) Ga (b) Al (c) Cl (d) Si
12. How many groups present in the modern periodic table  
(a) 16 (b) 17 (c) 18 (d) 20
13. How many periods present in the modern periodic table  
(a) 6 (b) 7 (c) 8 (d) 9
14. How many electrons present in f sub-shell  
(a) 2 (b) 6 (c) 10 (d) 14
15. Which one is Neutral charge particle  
(a) Proton (b) Positron (c) Neutron  
(d) Electron.
16. Anode rays are \_\_\_\_\_ charged particle  
(a) +ve (b) -ve (c) Neutral (d) zero

17. Non-metals are low density than Metals due to  
(a) weak force (b) strong force  
(c) repulsion (d) attraction.
18. Valency means what.  
(a) outermost electron (b) Inner electron  
(c) middle electron (d) loss of electrons.
19. Salt solution can be separated by  
(a) Filtration (b) Sublimation  
(c) Distillation (d) TLC method
20. Alloys are the  
(a) Gases (b) liquids (c) Solids  
(d) All of these
21.  $\text{Na}^+$  means what  
(a) loss of electrons (b) gaining of electrons  
(c) both (a) & (b) (d) None of these
22. Which one Reaction is Reduction.  
(a)  $\text{K} \rightarrow \text{K}^+ + \text{e}^-$  (b)  $\text{Cl} + \text{e}^- \rightarrow \text{Cl}^-$   
(c)  $\text{CH}_4 + \frac{1}{2}\text{O}_2 \rightarrow \text{CH}_3\text{OH}$  (d)  $\text{Na} + \text{Cl} \rightarrow \text{NaCl}$ .

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23. Noble gases are

(a) Stable (b) unstable (c) Isomorphous (d) half filled electrons.

24. Molecular mass of  $\text{CO}_2$

(a) 40g (b) 41g (c) 44g (d) 42g

25. Mass number contains

(a) proton + neutron (b) proton + proton (c) electron + electron (d) neutron + electron

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